

Mole Dispersion Unit Analysis & TIME RATE of CHANGE

MECH-322 Fluid Mechanics

Dr. K. J. Berry

Considering exit stream #3, the chemical molar dispersion [moles/sec] is measured as:

$$\psi(t, x, y) \left[\frac{\text{moles}}{s} \right] = 3.5e^{-5t} + 20.2x^2 - 0.3y^3$$

$$\vec{V}_3 = -V_3 \cos(\theta_3) \hat{i} + V_3 \sin(\theta_3) \hat{j}$$

What is the parametric form of the TIME RATE OF CHANGE of the chemical molar dispersion and VERIFY UNITS:

TIME RATE of CHANGE or TOTAL/MATERIAL DERIVATIVE

$$\frac{D\psi(t, x, y) \left[\frac{\text{moles}}{s} \right]}{Dt[s]} \left[\frac{\text{moles}}{s^2} \right] = \frac{\partial \psi}{\partial t} + V_x \frac{\partial \psi}{\partial x} + V_y \frac{\partial \psi}{\partial y}$$

UNIT ANALYSIS

$$\psi(t, x, y) \left[\frac{\text{moles}}{s} \right] = 3.5[\] e^{-5[\]t} + 20.2[\]x^2 - 0.3[\]y^3$$

"3.5"

$$\left[\frac{\text{moles}}{s} \right] = 3.5[\] \rightarrow [\] = \frac{\text{moles}}{s}$$

"-5"

$$[1] = -5[\] [s] \rightarrow [\] = \frac{1}{s}$$

"20.2"

$$\left[\frac{\text{moles}}{s} \right] = 20.2[\] [m^2] \rightarrow [\] = \frac{\frac{\text{moles}}{s}}{m^2} = \frac{\text{moles}}{s \cdot m^2}$$

"0.3"

$$\left[\frac{\text{moles}}{s} \right] = -0.3[\] [m^3] \rightarrow [\] = \frac{\frac{\text{moles}}{s}}{m^3} = \frac{\text{moles}}{s \cdot m^3}$$

$$\frac{D\psi(t, x, y) \left[\frac{\text{moles}}{s} \right]}{Dt[s]} \left[\frac{\text{moles}}{s^2} \right] = \frac{\overbrace{\frac{\partial \phi}{\partial t}}^{(-5)3.5e^{-5t}}}{\partial t} + \overbrace{\frac{\partial \phi}{\partial x}}^{-V_3 \cos \theta_3 \cdot 20.2 \cdot 2x} + \overbrace{\frac{\partial \phi}{\partial y}}^{V_3 \sin \theta_3 \cdot (-0.3 \cdot 3y^2)}$$

$$\psi(t, x, y) \left[\frac{\text{moles}}{s} \right] = 3.5 \square e^{-5\square t} + 20.2 \square x^2 - 0.3 \square y^3$$

"3.5"

$$\left[\frac{\text{moles}}{s} \right] = 3.5 \square \rightarrow \square = \frac{\text{moles}}{s}$$

"-5"

$$[1] = -5 \square [s] \rightarrow \square = \frac{1}{s}$$

"20.2"

$$\left[\frac{\text{moles}}{s} \right] = 20.2 \square [m^2] \rightarrow \square = \frac{\text{moles}}{s \cdot m^2} = \frac{\text{moles}}{s \cdot m^2}$$

"0.3"

$$\left[\frac{\text{moles}}{s} \right] = -0.3 \square [m^3] \rightarrow \square = \frac{\text{moles}}{s \cdot m^3} = \frac{\text{moles}}{s \cdot m^3}$$

$$= \overbrace{\left(-5 \left[\frac{1}{s} \right] \right) 3.5 \left[\frac{\text{moles}}{s} \right] e^{-5t}}^{\left[\frac{\text{moles}}{s^2} \right]} + \overbrace{\left(-V_3 \left[\frac{m}{s} \right] \cos \theta_3 \right) \left(20.2 \left[\frac{\text{moles}}{s \cdot m^2} \right] \cdot 2x [m] \right)}^{\left[\frac{\text{moles}}{s^2} \right]} + \overbrace{\left(V_3 \left[\frac{m}{s} \right] \sin \theta_3 \right) \left(-0.3 \left[\frac{\text{moles}}{s \cdot m^3} \right] \cdot 3y^2 [m^2] \right)}^{\left[\frac{\text{moles}}{s^2} \right]}$$